

# Acids and Bases (pH+pOH=14)

## Acids:

H<sup>+</sup> present; donate H<sup>+</sup> protons

Turn litmus paper red

pH less than 7

$$\text{pH} = -\log[\text{H}^+]$$

$$[\text{H}^+] = 10^{-\text{pH}}$$

## Bases:

OH<sup>-</sup> present; accept H<sup>+</sup> protons

Turn litmus paper blue

pH above 7

$$\text{pOH} = -\log[\text{OH}^-]$$

$$[\text{OH}^-] = 10^{-\text{pOH}}$$