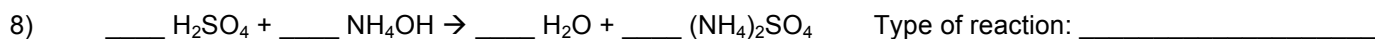
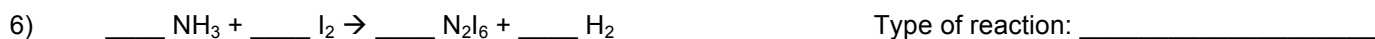
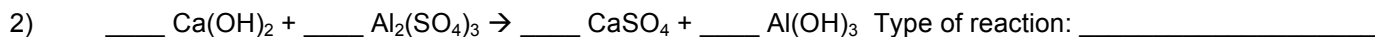


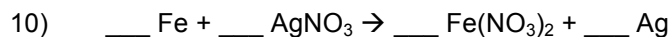
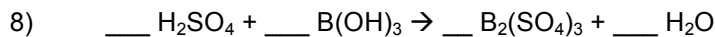
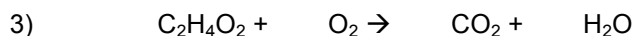
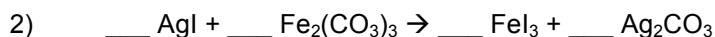
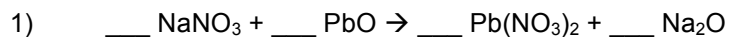
Types of Reactions Worksheet THEN Balancing!

First, begin by telling which type of reaction is taking place. Then go back and balance the following equations: To practice balancing, you may use the Phet Lab online. When finished, check your answers.



Balancing Equations Practice Worksheet

Balance the following equations:



Another Balancing Equations Sheet!

Balance these equations!

- 1) ___ AlBr_3 + ___ K \rightarrow ___ KBr + ___ Al
- 2) ___ FeO + ___ PdF_2 \rightarrow ___ FeF_2 + ___ PdO
- 3) ___ P_4 + ___ Br_2 \rightarrow ___ PBr_3
- 4) ___ LiCl + ___ Br_2 \rightarrow ___ LiBr + ___ Cl_2
- 5) ___ PbBr_2 + ___ HCl \rightarrow ___ HBr + ___ PbCl_2
- 6) ___ CoBr_3 + ___ CaSO_4 \rightarrow ___ CaBr_2 + ___ $\text{Co}_2(\text{SO}_4)_3$
- 7) ___ Na_3P + ___ CaF_2 \rightarrow ___ NaF + ___ Ca_3P_2
- 8) ___ Mn + ___ HI \rightarrow ___ H_2 + ___ MnI_3
- 9) ___ Li_3PO_4 + ___ NaBr \rightarrow ___ Na_3PO_4 + ___ LiBr
- 10) ___ CaF_2 + ___ Li_2SO_4 \rightarrow ___ CaSO_4 + ___ LiF
- 11) ___ HBr + ___ $\text{Mg}(\text{OH})_2$ \rightarrow ___ MgBr_2 + ___ H_2O
- 12) ___ LiNO_3 + ___ CaBr_2 \rightarrow ___ $\text{Ca}(\text{NO}_3)_2$ + ___ LiBr
- 13) ___ AgNO_3 + ___ Li \rightarrow ___ LiNO_3 + ___ Ag
- 14) ___ $\text{Si}(\text{OH})_4$ + ___ NaBr \rightarrow ___ SiBr_4 + ___ NaOH
- 15) ___ NaCN + ___ CuCO_3 \rightarrow ___ Na_2CO_3 + ___ $\text{Cu}(\text{CN})_2$

Word Equations Worksheet

Write and balance the following chemical equations.

- 1) When dissolved beryllium chloride reacts with dissolved silver nitrate in water, aqueous beryllium nitrate and silver chloride powder are made.
- 2) When isopropanol ($\text{C}_3\text{H}_8\text{O}$) burns in oxygen, carbon dioxide, water, and heat are produced.
- 3) When dissolved sodium hydroxide reacts with sulfuric acid (H_2SO_4), aqueous sodium sulfate, water, and heat are formed.
- 4) When fluorine gas is put into contact with calcium metal at high temperatures, calcium fluoride powder is created in an exothermic reaction.
- 5) When sodium metal reacts with iron (II) chloride, iron metal and sodium chloride are formed.