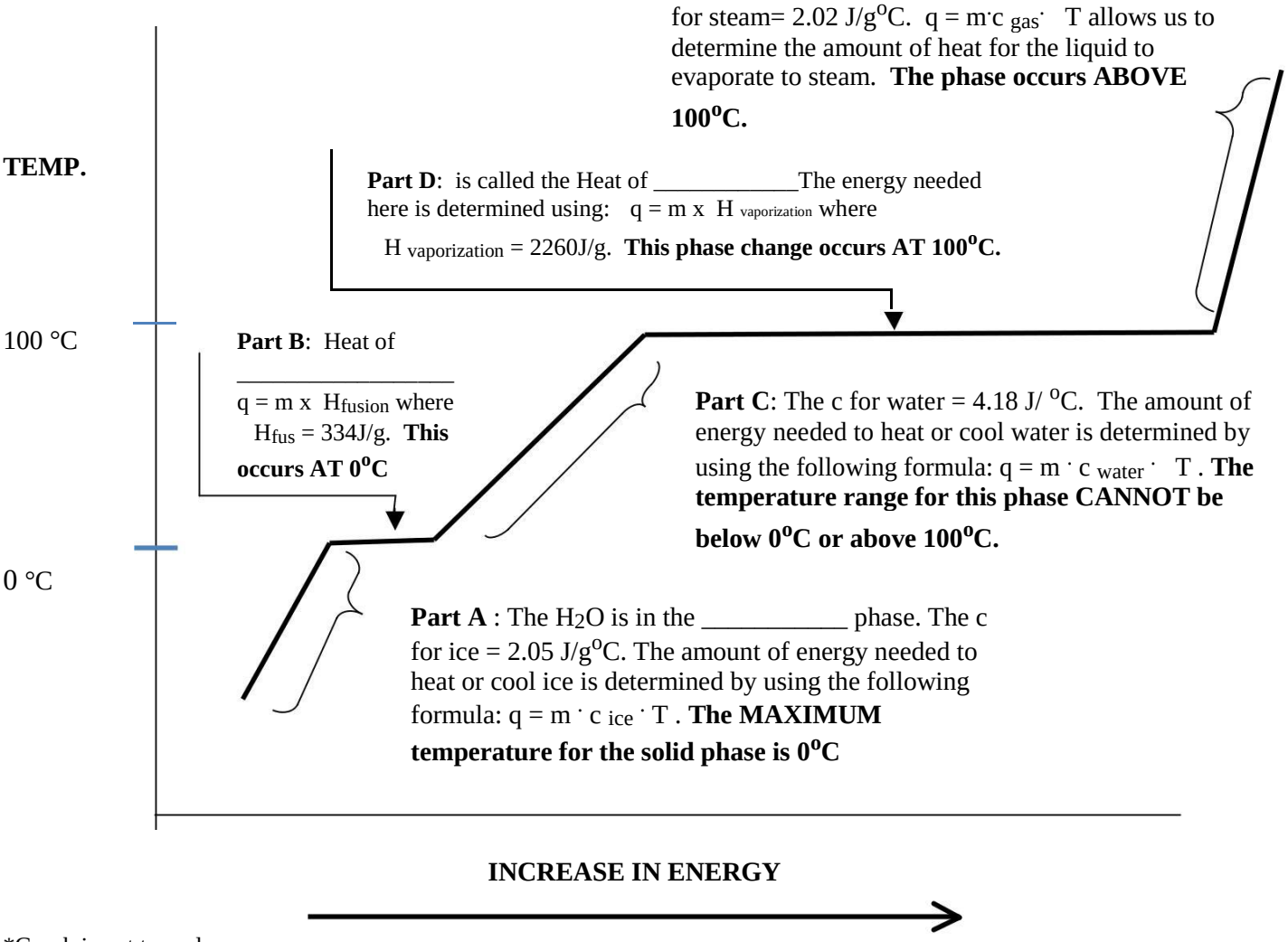
Name _____



*Graph is not to scale

How much heat does it take to heat 12 g of ice at – 6 °C to 25 °C water? Round to a whole number.

Because the ice is BELOW 0°C you must start with the solid phase.

q_A=

q_B=

q_C=

The amount of heat required to melt ice for it to become water goes through three stages. $q = q_A + q_B + q_C$

q=

How much heat does it take to heat 35 g of ice at 0 °C to steam at 150 °C? Round to a whole number.

Because the ice is AT 0°C you can start with the phase change from solid to liquid.

$q_B =$

$q_C =$

$q_D =$

$q_E =$

The amount of heat required to melt ice for it to become steam goes through four stages.

$q = q_B + q_C + q_D + q_E$

$q =$